

## NAMING INORGANIC ACIDS

BINARY ACIDS – start with hydrogen and end with a single non-metal element.

- Acid name ends in –ic.
- HCl – hydrogen chloride becomes **hydrochloric acid**
- HF – hydrogen fluoride becomes **hydrofluoric acid**
- HBr – hydrogen bromide becomes **hydrobromic acid**

OXYACIDS – Start with hydrogen and end with a polyatomic ion with OXYGEN in it.

- If polyatomic ion ends in –ate, acid ends in –ic.

EX.  $\text{HNO}_3$  – hydrogen **nitrate**, becomes **nitric acid**.

EX.  $\text{H}_2\text{SO}_4$  – hydrogen **sulfate**, becomes **sulfuric acid**

- If polyatomic ion ends in –ite, acid ends in –ous.
- EX.  $\text{H}_2\text{SO}_3$  – hydrogen **sulfite**, becomes **sulfurous acid**.
- EX.  $\text{HNO}_2$  – hydrogen **nitrite**, becomes **perchloric acid**.